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SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the

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theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂ according to the following equation: $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$

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CHAPTER 9 REVIEW. Stoichiometry. MIXED REVIEW. SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation: $\text{C}_3\text{H}_4(\text{g}) + x \cdot \text{O}_2(\text{g}) \rightarrow 3\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$ a. What is the value of the coefficient . x. in this equation? b. What is the molar mass of C₃H₄? c. How many moles are in an 8.0 g sample of C₃H₄? 2. a. What is meant by . ideal conditions

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space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the ...

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MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a. What is the value of the coefficient x in this equation? b. What is the molar mass of C_3H_4 ? c. What is the mole ratio of O_2 to H_2O in the above equation? d.

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