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## **Answer Key Chapter 22 Chemistry**

Answers to Chapter 22 Study Questions. 1. 2.

3. 5. should be

2-methylbutane;

counting from wrong

end. should be

2-methylbutane;

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longest chain is 4  
carbons, not 3 carbons;  
2-ethyl is impossible. 6.  
7.

## **Answers to Chapter 22 Study Questions - Chemistry LibreTexts**

22.1: Alkanes:

Saturated

Hydrocarbons

Hydrocarbons : -

compounds that

contains hydrogen and

carbon atoms. - it may

contain oxygen,

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nitrogen and other  
halogen atoms.

## **Chapter 22 Organic and Biological Molecules Notes (answers)**

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**Chapter 22 ...**

Modern Chemistry 183

Organic Chemistry

CHAPTER 22 REVIEW

Organic Chemistry

SECTION 3 SHORT

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ANSWER Answer the  
following questions in

the space provided. 1.

Match the structural  
formulas on the right  
to the family name on  
the left. \_\_\_\_\_ aldehyde

\_\_\_\_\_ ketone \_\_\_\_\_

carboxylic acid \_\_\_\_\_

amine \_\_\_\_\_ ester \_\_\_\_\_

alkene (a) (b) (e) (c) (d)

## **CHAPTER 22 REVIEW** **Organic Chemistry**

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Chemistry Eventually,

you will definitely



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discover a additional  
experience and  
realization by spending  
more cash. still when?

get you admit that you  
require to get those  
every needs similar to  
having significantly  
cash? Why dont you  
attempt to acquire  
something basic in the  
beginning?

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## Chemistry Study 22CHAPTER Study Guide Key Concepts

### 22.1 Hydrocarbons •

Because carbon has four valence electrons, carbon atoms always form four covalent bonds. • The carbon atoms in an alkane can be arranged in a straight chain or in a chain that has branches. • Molecules of hydrocarbons, such as alkanes, are nonpolar molecules.

### 22.2 Unsaturated

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## Chemistry Study Guide For Content

Hydrocarbons

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Example (answers may  
vary): For C-14, the

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half-life is 5770 years.  
A 10-g sample of C-14  
would contain 5 g of  
C-14 after 5770 years;  
a 0.20-g sample of  
C-14 would contain  
0.10 g after 5770  
years. 33 .

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1 . The instantaneous  
rate is the rate of a  
reaction at any  
particular point in time,  
a period of time that is

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so short that the  
concentrations of  
reacta

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General Properties of  
the Metalloids; 18.4  
Structure and General  
Properties of the  
Nonmetals; 18.5

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Occurrence,  
Preparation, and  
Compounds of  
Hydrogen; 18.6

Occurrence,  
Preparation, and  
Properties of  
Carbonates; 18.7

Occurrence,  
Preparation, and  
Properties of Nitrogen

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The Substituted  
Hydrocarbons and

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2 Measurements and  
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## Chemistry Study 2 REVIEW

### Measurements and Calculations SECTION 1 SHORT ANSWER

Answer the following questions in the space provided. 1. Determine whether each of the following is an example of observation and data, a theory, a hypothesis, a control, or a model ...

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1 . (a) reduction; (b)  
oxidation; (c)  
oxidation; (d) reduction

3 . (a)  $F_2 + Ca^{2+} \rightarrow 2F^- + Ca^{2+}$  ;  $F_2 + Ca^{2+} \rightarrow 2F^- + Ca^{2+}$  ; (b)  $Cl_2 + 2Li \rightarrow 2LiCl + 2.$

While the reduction potential of lithium would make it capable of protecting the other metals, this high potential is also indicative of how reactive lithium is; it would have a spontaneous reaction with most substances.

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(a) A straight horizontal line at 1.0; (b) When real gases are at low pressures and high temperatures, they behave close enough to ideal gases that they are approximated as such; however, in some cases, we see that at a high pressure and temperature, the ideal gas approximation

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breaks down and is significantly different from the pressure calculated by the ideal gas equation.

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CHAPTER 11 REVIEW  
Gases Class SHORT  
ANSWER Answer the  
following questions in  
the space provided. c c  
The molar mass of a  
gas at STP is the  
density of that gas (a)

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multiplied by the mass  
of 1 mol. (b) divided by  
the mass of 1 mol.  $nRT$   
(c) multiplied by 22.4  
L. (d) divided by 22.4 L.  
For the expression  $V =$   
(a) increasing P (b)  
decreasing T

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